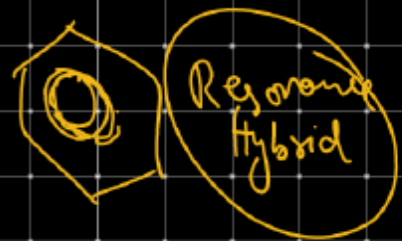
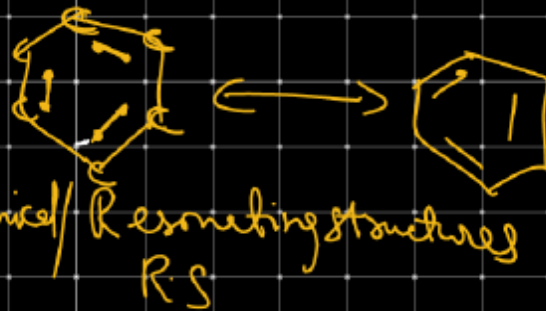
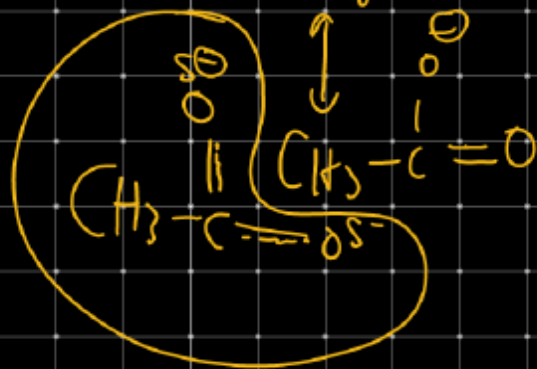
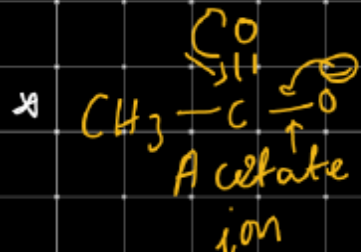
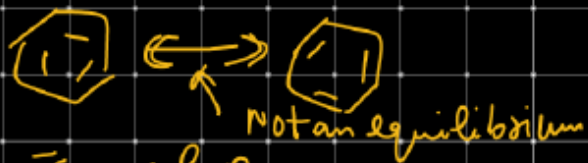


Resonance

* It is way to describe bonding in certain molecules in which bonding can't be explained using single Lewis Structure



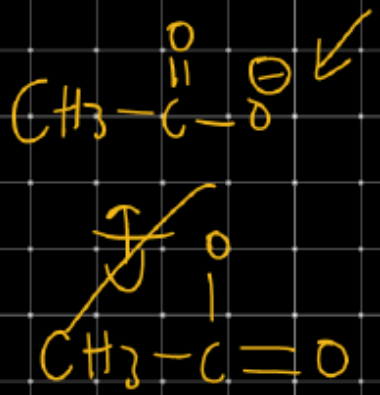
Characteristics



- * Delocalisation of π (π) e^- s or L.P
- * R.S are imaginary structure while R.H is real
- * More is the stability of R.S more will be the contribution in hybrid.
- * In resonance position of atoms doesn't change.
- * Resonance is a stabilising phenomenon. R.H is stable than all R.S.
- * It is permanent effect & stronger effect.
(Polarisation of π e's is greater than polarisation of σ e's)
- * For resonance participating orbitals should be parallel.



Condition for two structures to be called R-S

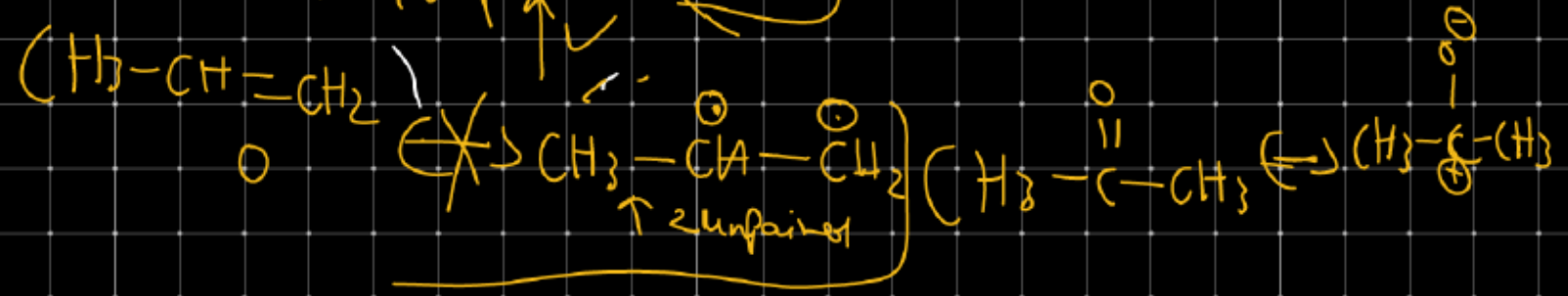


① Position of atoms should not change

② Total no. of e⁻s should remain same (Net charge should remain unchanged)

③ Total no. of paired & unpaired e⁻s should also remain same (σ, π, r.p) (F.R)

paired + unpaired = Total e⁻s



Case of Resonance

Case I π bond is in conjugation with vacant p-orbital or +ve charge

